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Chapter 3 - Stoichiometry, Formulas and Equations: Part 7 of 8

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Steve Zumdahl is the author of market-leading textbooks in introductory chemistry, general chemistry, honors-level chemistry, and high school chemistry. Recently retired from his long-time position as Director of Undergraduate Programs at the University of Illinois, he has received numerous awards for his contributions to chemical education.

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Step 1 of 1. The boiling point of water is not enough to break water molecules into hydrogen and oxygen. So, the bubble does not contain oxygen or hydrogen. Carbon dioxide may be present in small quantities in the water; however it is not at high enough quantities to form bubbles.

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Contains discussion, illustrations, and exercises aimed at overcoming common misconceptions; emphasizes on models prevails; and covers topics such as: chemical foundations, types of chemical reactions and solution stoichiometry, electrochemistry, and organic and biological molecules.

Written for calculus-inclusive general chemistry courses, Chemical Principles helps students develop chemical insight by showing the connections between fundamental chemical ideas and their applications. Unlike other texts, it begins with a detailed picture of the atom then builds toward chemistry ' s frontier, continually demonstrating how to solve problems, think about nature and matter, and visualize chemical concepts as working chemists do. It also offers an exceptional level of support to help students develop their mathematical and problem-solving skills. For the new edition, Chemical Principles now takes a modular approach, with coverage organized as a series of brief Topics within 13 major areas of focus, including a refresher on the fundamentals of chemistry and an online-only section on techniques.

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